## ICP: Density of Fluids

## Objectives:

1. Learn how to find the density of liquids and use your understanding to make a density column.
2. Use a density column to predict the density of a solid.

## Notes:

- A fluid is defined as any matter that is able to flow.

Density Column

## Density

1. What is density?

Density is a $\qquad$ of the amount of matter or $\qquad$ that can fit in a given
$\qquad$ . It is the ratio of a material's mass to its volume.

## 2. What is the formula for density?


3. What does density tell me about an object?

- Density is important because it affects whether objects will $\qquad$ or $\qquad$ .
- If density is less than $1.0 \mathrm{~g} / \mathrm{mL}$, the object $\qquad$ in water.
- If density is more than $1.0 \mathrm{~g} / \mathrm{mL}$, the object $\qquad$ in water.

4. Where would I see density in my everyday life?
$\qquad$ - you want the balloon to get off the ground.

* $\qquad$ - you don't want the ship to sink! People don't like to be on a sinking ship.
* $\qquad$ - you don't want to overload the truck!
$\qquad$ - you want to make sure you actually bought a gold necklace.

5. Density values chart

| Object | Density Value <br> $\left(\mathrm{g} / \mathrm{mL} \mathrm{or} \mathbf{~} / \mathrm{cm}^{\mathbf{3}}\right)$ | Sinks or Floats in <br> water? |
| :--- | :--- | :--- |
| Water $\left(4^{\circ} \mathrm{C}\right)$ |  |  |
| Ice water $\left(0^{\circ} \mathrm{C}\right)$ |  |  |
| Aluminum |  |  |
| Silver (solid) |  |  |
| Silver (liquid) |  |  |
| Gold |  |  |
| Iron |  |  |
| Gasoline |  |  |
| Baby Oil |  |  |

6. REMEMBER: the phase of matter depends on its $\qquad$ . The $\qquad$ the temperature, the farther the molecules are spread out and the faster they are moving. Therefore a $\qquad$ of one material will be denser than a $\qquad$ of the same material.
7. An exception to the rule is
$\qquad$ . Solid
$\square$ is less dense than liquid
$\qquad$ . That is why $\qquad$ floats in your drinks. molecules freeze into
$\qquad$ crystals that form a pattern that has a lot of empty space. The molecules in liquid $\qquad$ are more tightly packed. See the


Stable hydrogen bonds HEXAGONAL
 cubic illustration below.
8. Practice Problems
9. Clay Practice: A review of triple-beam balances and displacement method

| Clay Mass (g) | Volume (mL) | Density (g/mL) | Density (g/mL) <br> (rounded to nearest whole <br> number) | Sinks or Floats? |
| :--- | :--- | :--- | :--- | :--- |
| 10 |  |  |  |  |
| 20 |  |  |  |  |
| 30 |  |  |  |  |

