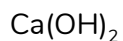
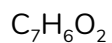


HOMEWORK: GROUP/SELF PRACTICE: Solve each problem below. When you have finished EVERY problem, please come to the teacher's lab table and check your work.

► **Molar Mass. Calculate the molar mass of each compound below.**



► **Solve for the unknown.**

1. How many moles are 1.20×10^{25} atoms of phosphorus?

2. How many moles of argon atoms are present in 11.2 L of argon gas at STP?

3. How many molecules are in 0.400 moles of N_2O_5 ?

4. What is the mass of 5 moles of Fe_2O_3 ?

5. Find the number of moles of argon in 452 g of argon.

6. Find the grams in 1.26×10^{-4} mol of $\text{HC}_2\text{H}_3\text{O}_2$.

7. Find the mass in 2.6 mol of lithium bromide (LiBr).

8. What is the volume of 0.05 mol of neon gas at STP?

9. How many molecules are there in 17.2 moles of glucose, $\text{C}_6\text{H}_{12}\text{O}_6$?

10. How many atoms are in 0.750 moles of zinc?